

Student name:

Student number:

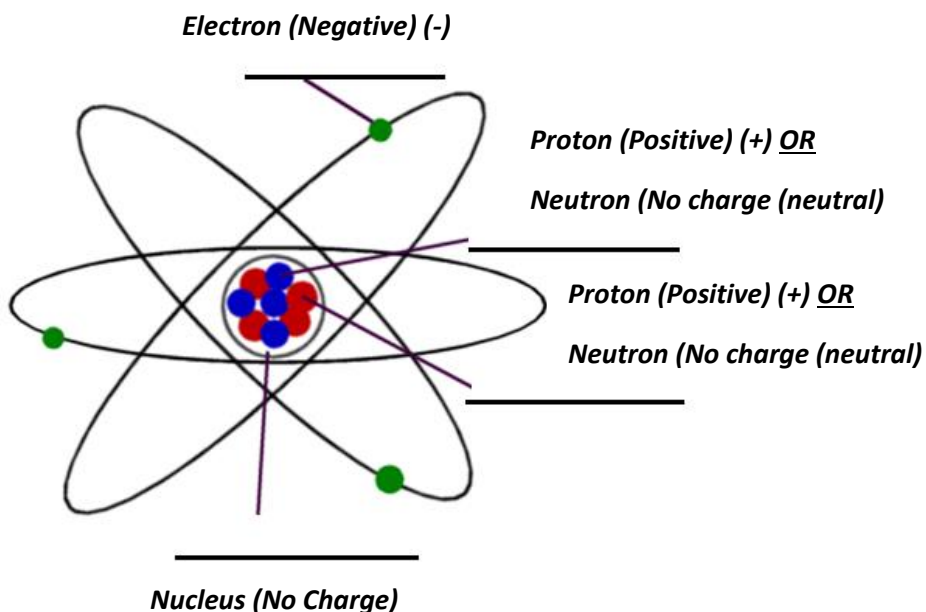
All About Atoms

Activity 1: Listen to the presentation and fill in the missing information below.

Fill in the blanks:

Atoms make up **everything** in the universe. The three parts that make up an atom are **protons**, neutrons and **electrons**. Protons and **neutrons** are inside the **nucleus**. **Electrons** go around the atom.

Label the diagram with the correct name, include the charges of the particle.



Fill in the blanks:

Atoms are so small that we can't see them. They are made of mostly **empty space**, but atoms look **solid** because they are moving so fast.

The number of **protons** in an atom's nucleus change what the atom does and looks like.

Different atoms are called **elements**. All of the elements on the Earth are on the **Periodic Table**.

Student name:

Student number:

The atomic number is equal to the number of protons an atom has. The atomic mass is how heavy an atom is, this is almost equal to the mass of the protons and neutrons of an atom.

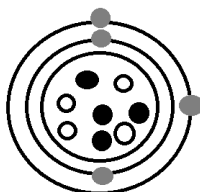
Activity 2: Answer the questions below.

1. What is the atomic number and atomic mass of the element Sulfur? Why does it have that atomic number?



The atomic number of Sulfur is 16. The atomic mass of Sulfur is 32.066. It has that atomic number because it has 16 protons in its nucleus.

2. Using the Periodic Table, find which element this diagram is representing.



This diagram is representing Beryllium (Be).

3. What is an ion?

An ion is when an atom has more or less electrons than protons.

4. Circle the atom(s) below that is/are (an) ion(s).

